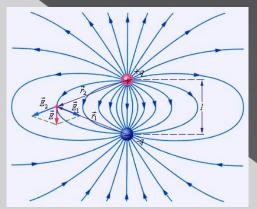
ELECTRICAL POTENTIALS OF THE CELL

Electric field. Coulomb's law.

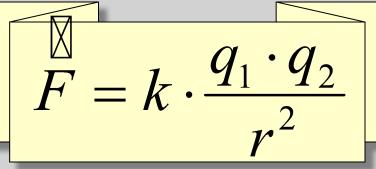
Electric field is the kind of matter by means of which the charged objects interact with each other.

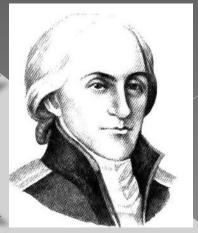


Coulomb law defines the value of electric force Facting between the two electric

charges q_1 and q_2

Here *k* is the constant determined by the measure units chosen and is the distance between the charges.



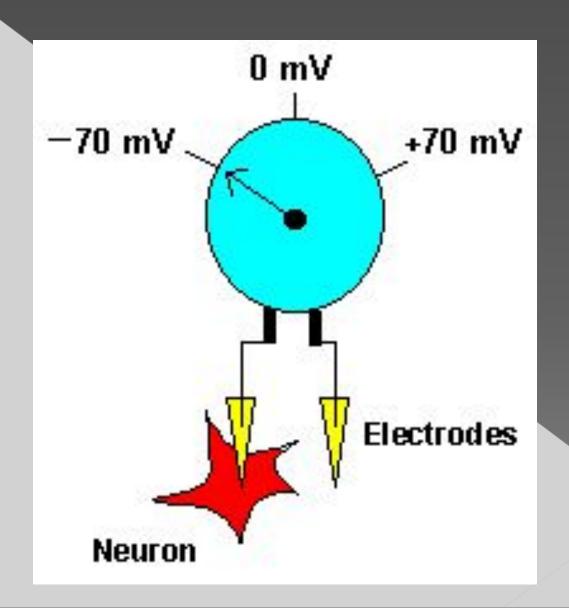


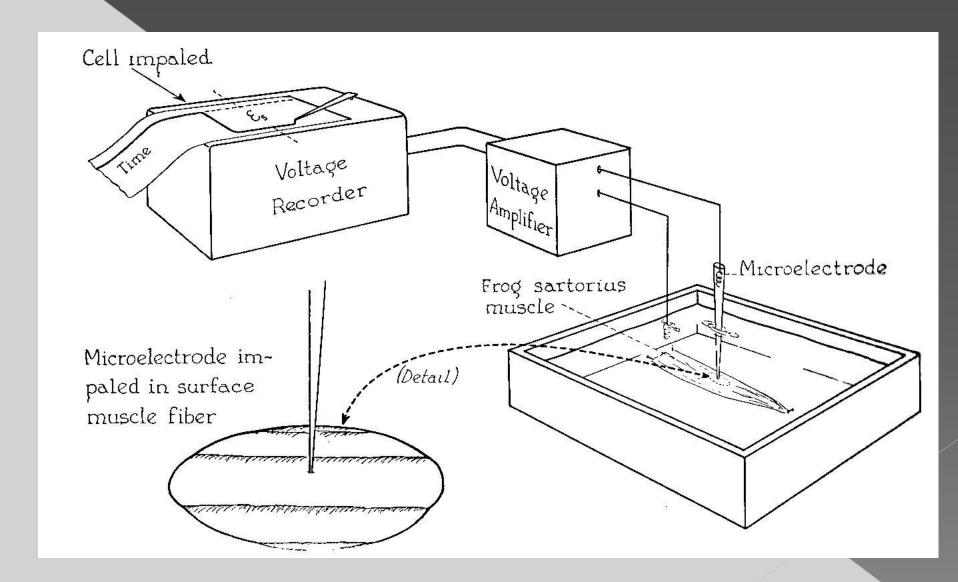
The electric field intensity *E* at any point is defined as the electric force *F* that would be exerted on the unit positive charge *q* placed at this point

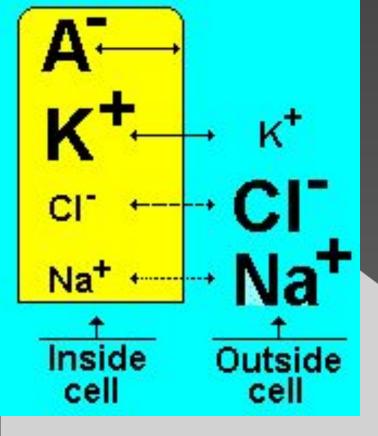
$$E = \frac{F}{q}$$

The electric potential ϕ (voltage) at any point equals the electrostatic potential energy W_p that the unit positive charge q would have at this point

$$\varphi = \frac{W_p}{q}$$







INSIDE

OUTSIDE

K+ A-	Na+ CI-
A- A- K+	Na+ CI-
CI- CI- K+	CI- Na+ Na+
A- Na+ K+	Na+ CI-
K+ A-	K+ Na+ CI-
A- K+	CI- Na+ ^{Na+}
K+ CI-	CI- CI- Na+ CI-

With no channels, there is no diffusion force.

Kind of ions		Cell protoplasm	Extracellular fluid
	Sodium	12	145
Cations	Potassium	155	4
	Others	040	5
Anions	Chloride	4	120
	Bicarbonate	8	27
	Others (A)	155	7

$$E = -\frac{RT}{zF} \ln \frac{P_K[K^+]i + P_{Na}[Na^+]i + P_{Cl}[Cl^-]o}{P_K[K^+]o + P_{Na}[Na^+]o + P_{Cl}[Cl^-]i}$$

Goldman-Hogkin equation for membrane potential

Nernst equation for membrane potential

$$E = -\frac{RT}{zF} \ln \frac{[K^+]i}{[K^+]o}$$

$$E_{m} = \frac{RT}{\mathbf{z}F} \ln \left(\frac{P_{K}[K^{+}]_{out} + P_{Na}[Na^{+}]_{out} + P_{Cl}[Cl^{-}]_{in}}{P_{K}[K^{+}]_{in} + P_{Na}[Na^{+}]_{in} + P_{Cl}[Cl^{-}]_{out}} \right)$$

